



Introductory Chemistry

Chemistry 1010 Hong Kong •
Joe SUEN and Doug Harris
Spring 2008 Course Schedule
3 credits

Date	Thursday, 14:00 – 16:30
10 th Jan 2008	Ch 1-2
17 th Jan 2008	Ch 3-4
24th Jan 2008	Exam 1
31 st Jan 2008	Ch 5-6
7th Feb 2008	Holiday
14 th Feb 2008	Ch 6-7
21st Feb 2008	Exam 2
28 th Feb 2008	Ch 8
6 th Mar 2008	Ch 9
13th Mar 2008	Exam 3
20 th Mar 2008	Ch10
27 th Mar 2008	Ch11
3 rd Apr 2008	Ch12
10th Apr 2008	Exam 4
17 th Apr 2008	Ch13
24th Apr 2008	Final Exam

Exam Number	Date	Chapters Included
1	24 th Jan 2008	1 – 3
2	21 st Feb 2008	4 – 6
3	13 th Mar 2008	7 – 9
4	10 th Apr 2008	10 – 12
Final	24 th Apr 2008	1 – 13

Facilitator Email Mr. Joe SUEN Email: suenwai99@hotmail.com
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Materials

John Suchocki, *Conceptual Chemistry*, 3rd edition (custom-published text), Pearson, 2007.
 Scientific Calculator (no cell phone calculators)
 Course web site: <http://www.chem.usu.edu/~harrisd/>

Coursework

Examinations, 4 @ 100..... 400
 Final Exam, comprehensive @ 100..... 100
 TOTAL (drop lowest exam score)..... 400

Grades

100% - 92%	A
91% - 88%	A-
87% - 85%	B+
84% - 81%	B
80% - 77%	B-
76% - 73%	C+
72% - 64%	C
63% - 60%	C-
59% - 57%	D+
56% - 50%	D

Note: Scores rounded to nearest one's place (91.4% = 91% and 91.5% = 92%).
The instructors reserve the right to lower these cutoff scores.

Policies and Procedures

1. The administration of Chemistry 1010 will adhere strictly to the policies outlined in the USU Spring 2008 Semester Schedule of Classes.
2. Qualified students with disabilities may be eligible for reasonable accommodations. All accommodations are coordinated through the Disability Resource Center (DRC) in Room 101 of the University Inn, 797-2444 voice, 797-0740 TTY, or toll free at 1-800-259-2966. Please contact the DRC as early in the semester as possible. Alternate format materials (Braille, large print or digital) are available with advance notice.
3. Except for school-excused absences, exams will not be rescheduled. All exam rescheduling requests for non-school excused absences will be directly referred to this policy without any further discussion.
4. Keep in mind that the practice exam serves as an assessment of your understanding of concepts presented in lecture. Hopefully you will be diligent about following the suggested study plan outlined at the beginning of the course. Exam questions may be the same or similar to the practice exam problems but may also be completely different. Although exam questions may be completely different from the practice exam problems, all exam questions will focus on the concepts discussed in lecture.
5. Scantrons will be provided by the instructor.
6. When taking the exam, be sure to answer the problem and immediately fill out the corresponding scantron bubble. Avoid waiting to fill out your scantron sheet when finished with your exam.
7. Double check your scantron sheet before turning it in. Make sure that all of your answers have been entered the way you want them to appear on your scantron. Once an exam scantron is submitted, it may not be retrieved in order to make additions and/or changes.
8. Please arrive early to take the exam. Exams and scantron sheets will not be handed out after the first completed exam scantron sheet has been submitted. All requests for an exam and scantron sheet after the first completed exam scantron sheet has been submitted will be directly referred to this policy without further discussion.
9. Although class attendance will not be officially taken, it will be absolutely essential that every effort is made in attending each lecture. All students will be held responsible for lecture material, worked problems, and/or course announcements that are presented in lecture.
10. If you choose to complete an optional extra-credit molecular modeling exercise, one percentage point (1%) will be added to your final grade percentage. This is helpful to those students who end up with a final borderline grade percentage. The extra-credit submission deadline will be at 14:00, Thursday, March 13th when we meet for the third exam. Further information will be given in class on Thursday, February 14th regarding the specific details in producing the extra-credit assignment.

Main Course Objective and Assessment

1. The course will present chemistry *conceptually*, focusing on the concepts of chemistry with little emphasis on calculations. This presentation will hopefully improve each student's learning skills and assist in developing better thinking abilities (text page xix).
2. Lecture learning checks will be used as a means of assessing student comprehension. These student-centered learning strategies have previously proven successful in this chemistry course.

Some Learning Objectives:

- Become familiar with the basic physical quantities including mass, volume, energy, temperature, and density.
- Understand the fundamental concepts and language of chemistry including physical properties, chemical properties, elements, mixtures, compounds, and atomic structure.
- Understand how elements are organized in the periodic table.
- Understand radioactivity, three major radioactive products, and half-life of a radioactive isotope.
- Explore two types (ionic and covalent) of chemical bonds.
- Given a covalent molecular formula, predict the molecular structure.
- Describe the various types of intermolecular interactions.
- Gain an understanding of the macroscopic consequences of water's molecular structure.
- Gain an understanding of the basics of chemical reactions.
- Explore acids and bases and the chemical reactions they undergo.
- Gain a basic understanding of organic compounds.
- Gain a basic understanding of biomolecules.