



Introductory Chemistry

Chemistry 1010 - 001 • Dr. Harris
 Fall 2006 Course Schedule
 8:30 – 9:20 a.m., MWF
 Widtsoe 007
 3 credits

Dates		MON	WED	FRI
August 28 th – 30 th	September 1 st	Introduction	1	1
September	4 th – 8 th	Holiday	1	2
September	11 th – 15 th	2	2	3
September	18 th – 22 nd	3	[E1] 8:30 a.m. Widtsoe 007	4
September	25 th – 29 th	4	4	5
October	2 nd – 6 th	5	5	6
October	9 th – 13 th	6	6 Extra Credit Information	[E2] 8:30 a.m. Widtsoe 007
October	16 th – 20 th	7	7	Holiday
October	23 rd – 27 th	8	8	8
October 30 th – 31 st	November 1 st – 3 rd	9	9	9
November	6 th – 10 th	[E3] 8:30 a.m. Widtsoe 007	10	10
November	13 th – 17 th	10	11	11
November	20 th – 24 th	11	Holiday	Holiday
November 27 th – 29 th	December 1 st	12	12	[E4] 8:30 a.m. Widtsoe 007
December	4 th – 8 th	13	13	13
December	11 th	[Final] 7:30 a.m. Widtsoe 007		

Exam Number	Date	Chapters Included
1	Wednesday, 20 th of September	1 – 3
2	Friday, 13 th of October	4 – 6
3	Monday, 6 th of November	7 – 9
4	Friday, 1 st of December	10 – 12
Final	Monday, 11 th of December	1 – 13

Dr. Doug Harris
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 E-mail: harrisd@cc.usu.edu
 Office Hours: 10:30 – 11:30 a.m. MW

Materials

John Suchocki, *Conceptual Chemistry*, 3rd edition (custom-published text), Benjamin Cummings, 2007.
 Scientific Calculator (no cell phone calculators)
 Course web site: <http://www.chem.usu.edu/~harrisd/>

Coursework

Examinations, 4 @ 100..... 400
 Final Exam, comprehensive @ 100..... 100
 TOTAL (drop lowest exam score)..... 500

Grades

100% - 92%	A
91% - 88%	A-
87% - 85%	B+
84% - 81%	B
80% - 77%	B-
76% - 73%	C+
72% - 64%	C
63% - 60%	C-
59% - 57%	D+
56% - 50%	D

Note: Scores rounded to nearest one's place (91.4% = 91% and 91.5% = 92%).
The instructor reserves the right to lower these cutoff scores.

Policies and Procedures

1. The administration of Chemistry 1010 will adhere strictly to the policies outlined in the USU Fall 2006 Semester Schedule of Classes.
2. Qualified students with disabilities may be eligible for reasonable accommodations. All accommodations are coordinated through the Disability Resource Center (DRC) in Room 101 of the University Inn, 797-2444 voice, 797-0740 TTY, or toll free at 1-800-259-2966. Please contact the DRC as early in the semester as possible. Alternate format materials (Braille, large print or digital) are available with advance notice.
3. Except for school-excused absences, exams will not be rescheduled.
4. Keep in mind that the practice exam serves as an assessment of your understanding of concepts presented in lecture. Hopefully you will be diligent about following the suggested study plan outlined at the beginning of the course. Exam questions may be the same or similar to the practice exam problems but may also be completely different. Although exam questions may be completely different from the practice exam problems, all exam questions will focus on the concepts discussed in lecture.
5. Scantrons will be provided by the instructor.
6. When taking the exam, be sure to answer the problem and immediately fill out the corresponding scantron bubble. Avoid waiting to fill out your scantron sheet when finished with your exam.
7. Double check your scantron sheet before turning it in. Make sure that all of your answers have been entered the way you want them to appear on your scantron. Once an exam scantron is submitted, it may not be retrieved in order to make additions and/or changes.
8. Please arrive early to take the exam. Exams and scantron sheets will not be handed out after the first completed exam scantron sheet has been submitted.
9. Exam cheating is a direct violation of the University Honor Code and will not be tolerated.
10. Although class attendance will not be officially taken, it will be absolutely essential that every effort is made in attending each lecture. All students will be held responsible for lecture material, worked problems, and/or course announcements that are presented in lecture.
11. If you choose to complete an optional extra-credit molecular modeling exercise, one percentage point (1%) will be added to your final grade percentage. This is helpful to those students who end up with a final borderline grade percentage. The extra-credit submission deadline will be at 8:30 a.m. Monday, November 6th when we meet to take the third exam. Further information will be given in class on Wednesday, October 11th regarding the specific details in producing the extra-credit assignment.

Main Course Objective and Assessment

1. The course will present chemistry *conceptually*, focusing on the concepts of chemistry with little emphasis on calculations. This presentation will hopefully improve each student's learning skills and assist in developing better thinking abilities (text page xix).
2. Use student centered learning strategies in the chemistry classroom.

Some Learning Objectives:

- Become familiar with the basic physical quantities including mass, volume, energy, temperature, and density.
- Understand the fundamental concepts and language of chemistry including physical properties, chemical properties, elements, mixtures, compounds, and atomic structure.
- Understand how elements are organized in the periodic table.
- Understand radioactivity, three major radioactive products, and half-life of a radioactive isotope.
- Explore two types (ionic and covalent) of chemical bonds.
- Given a covalent molecular formula, predict the molecular structure.
- Describe the various types of intermolecular interactions.
- Gain an understanding of the macroscopic consequences of water's molecular structure.
- Gain an understanding of the basics of chemical reactions.
- Explore acids and bases and the chemical reactions they undergo.
- Gain a basic understanding of organic compounds.
- Gain a basic understanding of biomolecules.